

Studies of Novel Copolymers for Deep-UV Photoresists. I. Synthesis and Properties of Poly(Styrene-*co*-Silicon-containing maleimide)

Wei-Jye SHU[†]

Department of Chemical and Material Engineering, Ta-Hwa Institute of Technology,
1, Ta-Hwa Road, Chung-Lin, Hsinchu 30743, Taiwan, ROC

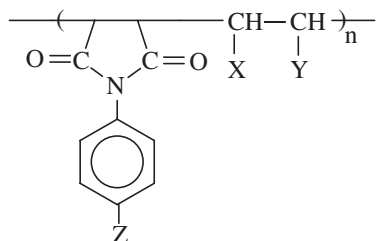
(Received December 7, 2005; Accepted May 22, 2006; Published July 28, 2006)

ABSTRACT: Soluble copolymers of trimethyl (4-(*N*-maleimido) phenoxy) silane (TMMS) with styrene-series monomers were synthesized by radical polymerization in toluene at 70 °C using 2, 2'-azobisisobutyronitrile (AIBN) as initiator. The comonomer reactivity ratios were calculated by the conventional Fineman-Ross and Kelen-Tüdös methods and a nonlinear least-squares Tidwell-Mortimer method. The glass transition temperatures (T_g s) and thermal degradation of copolymers were determined by differential scanning calorimetric (DSC) and thermo-gravimetric analysis (TGA) methods, respectively. The comonomer reactivity ratio of TMMS with styrene was changed from an alternating to middle type between alternating and ideal copolymerization for the side-chain of TMMS and the deactivators of styrene. The curves of T_g s versus the different compositions of the above synthesized styrene-maleimide copolymers matched the modified Johnston's equation as increasing the affect of weight ratio on alternating-segment, and exhibited an S-shaped curve of deviation in comparison with the Fox's equation. The maleimide-segments and styrene-segments within these copolymers were completely compatible and the thermal stability and flame retardancy of polystyrene could be enhanced simultaneously *via* the introduction of silicon-containing maleimide.

[doi:10.1295/polymj.PJ2005186]

KEY WORDS Poly(styrene-*co*-maleimide) / Copolymerization / Reactivity Ratio /

Maleimide polymers have good thermal stability with high glass transition temperature (T_g) due to their polar five membered imide ring structure. The functional polymaleimides made from various functionalized maleimide monomers also provide some desirable properties. *N*-aryl maleimide (AM) monomers, such as *N*-phenyl maleimide (PM), *N*-hydroxyphenyl maleimide (HPM) and halide-substituted *N*-hydroxyphenyl maleimide (XHPM), were usually copolymerized with ethylene or propylene series monomers to promote their heat and fire resistance. Many researches about the AM copolymers, such as the copolymers with styrene,¹⁻⁴ methyl methacrylate,⁵⁻⁷ acrylonitrile⁸ or vinyl acetate,⁹ have been reported.



These copolymers synthesized by introducing the rigid and high thermal stability of AM segment own the higher softening points (or T_g) and good fire resistance. Comparing the double bonds of AM and styrene comonomers with poor-electron and rich-electron

density, respectively, a nearby complete alternating copolymer can be formed by free radical copolymerization. The copolymerization behavior of truly alternating styrene-maleic anhydride copolymer has already been reported by many researches.¹⁰⁻¹³ But, the silicon-containing AM systems are still rarely seen in literature.

Up to now, applications of *N*-substituted maleimide polymers include photoresist with high T_g ,^{14,15} flexibilizer for thermosetting polymers,¹⁶⁻¹⁸ non-linear polymer with high T_g and flame retardant¹⁹ etc. Some *N*-protected polymaleimides have been investigated for their applicability as thermally-stable resist materials in the deep ultraviolet (DUV) region based on the chemical amplification concept.²⁰⁻²³ The photoresist film should possess high resolution, high sensitivity, high etching resistance, good thermal stability, and high adhesion for the lithographic processes. Incorporating silicon in the polymaleimide structure could advantageously provide good dry-etch resistance and enhanced adhesion to substrates.^{24,25}

Our team has already designed a series of different silicon-containing side chains introducing into *N*-aryl maleimides with high thermal stability, which could be polymerized by free radical initiator to form the soluble homopolymer.²⁶ These silicon-containing maleimides will then be combined with styrene-series

[†]To whom correspondence should be addressed (Tel: +886-3-5923551, Fax: +886-3-5927310, E-mail: shu@thit.edu.tw).

monomers in order to obtain various copolymers. They might be useful in the application of deep-UV photoresist for the acid-degradable side-chain within maleimides and styrenes. The copolymerization of a series of silicon-containing maleimide monomers with styrene have also been discussed simultaneously. In this paper we will discuss the reactivity ratios of four styrene-series monomers/silicon-containing *N*-aryl maleimide pairs in copolymerization to study their copolymerized behaviors. The T_g s and thermal stabilities of the above synthesized copolymers are also discussed. The possibility of these copolymers applied as photoresist materials will then be confirmed further more.

EXPERIMENTAL

Materials

Styrene (St) was distilled under reduced pressure after removal of the inhibitor by the usual method and kept cold before use. 4-Acetoxystyrene (A-St), *t*-Butyl 4-vinyl-phenyl carbonate (B-St), and 4-chlorostyrene (C-St) were used without further purification. Trimethyl(4-(*N*-maleimido)phenoxy) silane (TMMS) were prepared by the method of our previous paper.²⁶ The synthesized products were dissolved in ethyl acetate and extracted by a 1% NaOH solution. The organic layer was isolated and the remains, dried with anhydrous magnesium sulfate, were then dissolved in ethyl acetate and recrystallized several times with *n*-hexane. The silicon-containing maleimide monomers obtained were dried in vacuum. The characterization of this monomer synthesized was analyzed by the ¹H, ¹³C, ²⁹Si Nuclear Magnetic Resonance (NMR) spectroscopy, Fourier Transform Infrared (FT-IR) reflection spectroscopy and elemental analysis, respectively.

The initiator, AIBN, was recrystallized from ethanol and dried in a desiccator. Its melting point analyzed by DSC was found to be 104.1 °C. Tetrahydrofuran (THF) was distilled after dehydration with sodium. *N,N*-Dimethylformamide (DMF) was dried by CaH₂ over night. The other solvents were purified by conventional methods.

Synthesis of Copolymer

A flask was charged with various ratios of styrene and maleimide monomers (MI) obtained as shown in Table I, AIBN as initiator (5 mmol/L) and toluene as solvent. The free radical copolymerization was carried out in a water bath thermostat at 70 °C and the mixture was refluxed for a time in which the conversion of copolymerization was restricted to less than 10%. The solvent was removed under reduced pressure, and then the polymerization mixture was again dissolved in dichloromethane and poured into a large

Table I. Properties of copolymers prepared from styrene monomers (M₁) and TMMS (M₂)

Monomer (M ₁)	F_1^a	Conversion (%)	f_1^a	M_w^b ($\times 10^4$)	M_n^b ($\times 10^4$)	PDI	
St	0.10	9.1	0.18	24.6	7.2	3.42	
	0.25	9.4	0.31	25.7	7.6	3.38	
	0.40	8.9	0.44	26.9	8.1	3.32	
	0.50	9.0	0.54	27.3	8.4	3.25	
	0.60	7.6	0.57	27.6	8.6	3.21	
	0.75	9.8	0.71	28.7	9.1	3.15	
	0.90	9.2	0.86	28.7	9.5	3.02	
	A-St	0.10	9.0	0.14	20.4	5.8	3.51
		0.25	8.8	0.28	21.7	6.3	3.45
0.40		9.6	0.42	23.2	6.8	3.41	
0.50		9.3	0.52	23.6	7.0	3.37	
0.60		9.4	0.64	24.6	7.4	3.32	
0.75		9.9	0.74	25.6	7.8	3.28	
0.90		9.2	0.88	26.4	8.2	3.22	
B-St		0.10	8.6	0.13	18.5	5.2	3.55
		0.25	8.5	0.27	19.5	5.6	3.49
	0.40	8.9	0.41	21.2	6.2	3.42	
	0.50	9.0	0.51	23.0	6.8	3.38	
	0.60	9.2	0.62	23.7	7.1	3.34	
	0.75	9.6	0.75	24.8	7.5	3.30	
	0.90	8.8	0.89	25.6	7.8	3.28	
	C-St	0.10	8.7	0.17	24.2	7.0	3.45
		0.25	8.7	0.29	25.2	7.4	3.41
0.40		9.2	0.43	26.6	7.9	3.37	
0.50		9.2	0.53	26.8	8.1	3.31	
0.60		9.0	0.56	27.2	8.4	3.24	
0.75		9.7	0.73	27.7	8.8	3.15	
0.90		9.5	0.87	28.0	9.1	3.08	

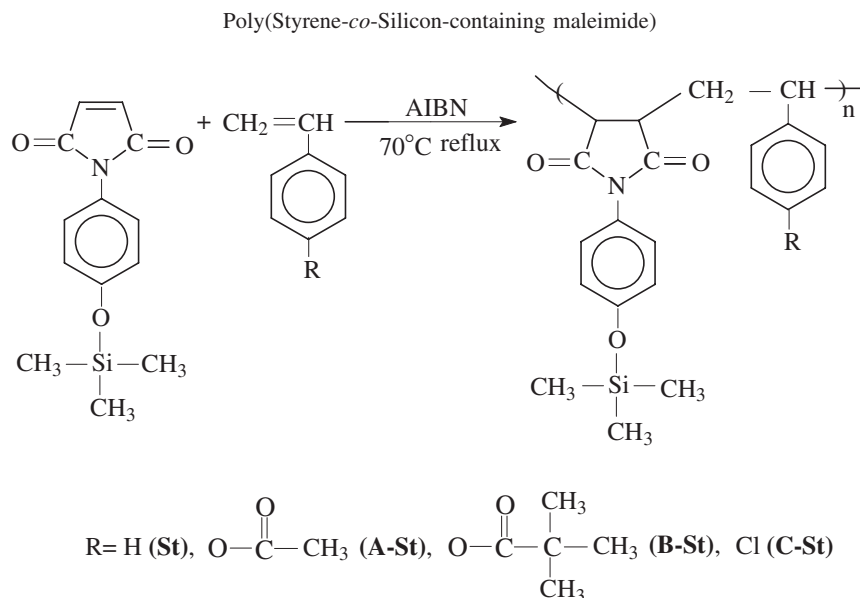
^a F_1 and f_1 are the mole fraction of M₁ in the feed and in the copolymers while f_1 are obtained from element analysis.

^bDistribution of molecular weight are tested by GPC with a concentration of 0.5 wt % in THF.

amount of methanol. The copolymer then was isolated by filtration and dried under vacuum at 60 °C.

Instrumentation

The IR-spectra and elemental analyses (C, H, N) of the St-MI copolymers were obtained by a Nicolet Omnic 3 Fourier Transform Infrared (FT-IR) reflection spectroscopy and a Heraeus CHN-O Rapid Analyzer, respectively. The molecular weight distributions of the St-MI copolymers were measured by a Waters Gel Permeation Chromatography (GPC, Waters 1515HPLC/2410RI-detector) at a flow rate of 1.0 mL/min with a sample concentration of 0.5 wt % in THF as effluent flow. The system was first calibrated using standard samples of polystyrene with the narrow distribution of molecular weight. The thermal properties of the synthesized copolymers were investigated by DSC(TA-Instruments DSC –Q10) at a heating rate of 10 °C/min in N₂, and TGA(TA-Instru-



Scheme 1. Synthesis of silicon-containing St-MI copolymer.

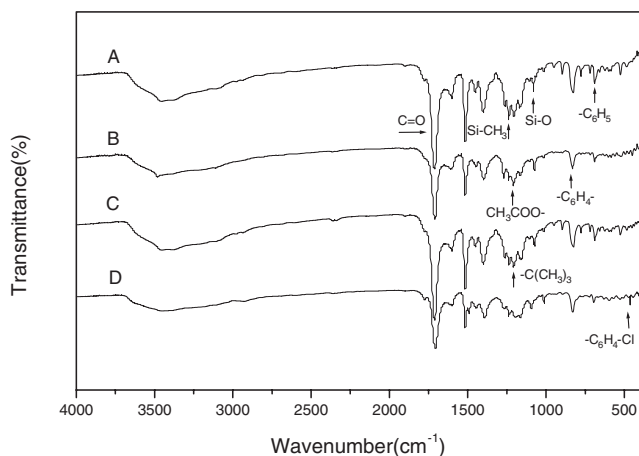


Figure 1. FT-IR spectra of A:St/MI(50%), B:A-St/MI(50%), C:B-St/MI(50%) and D:C-St/MI(50%) copolymer.

ments TGA –Q50) at a heating rate of 20 °C/min in air or N₂, respectively.

RESULTS AND DISCUSSION

Copolymerization and Copolymer Properties

The copolymers of styrene-series with silicon-containing *N*-phenyl maleimide were synthesized by radical polymerization in toluene solution using AIBN as initiator as shown in Scheme 1.

The FT-IR spectra of St/MI(50%), A-St/MI(50%), B-St/MI(50%) and C-St/MI(50%) copolymers, as shown in Figure 1, revealed the absorption bands $\nu(\text{cm}^{-1})$: 1240(Si–CH₃), 1080(–Si–O) and 1706(C=O) from TMMS; 700(mono-substitution of benzene) from styrene; 831(di-substitution of benzene) and 1210(CH₃COO–) from A-St; 1270(–C(CH₃)₃) from B-St and 463(–C₆H₄–Cl) from C-St conforming that all the

4 series copolymers had been synthesized successfully.

Basic properties of St-MI copolymers in different silicon contents and molecular weights obtained from elemental analysis and GPC analysis are given in Table I. The average number molecular weights (M_n) of each copolymer series were approximately in the range $5.2\text{--}9.5 \times 10^4$ (g/mol). The M_n gradually decreased, while the polydispersity index ($\text{PDI} = M_w/M_n$) increased along with the increasing maleimide monomers in feed. This might due to that the silicon-containing maleimide monomer could easily cause the chain transfer during copolymerization reaction. Besides, the copolymerization degree and average molecular weight might be affected by the side-chain effect and steric hinderance generated from maleimide and styrene segments. The sequence of intermolecular steric hinderance should be B-St > A-St > C-St > St. Since B-St series copolymers would form a larger steric hinderance during copolymerization, a lower polymerization degree and a lower average molecular weight would be obtained.

Generally speaking, the double bond of *N*-phenyl maleimide as a result of the electron-withdrawing effect between carbonyl groups on both sides showed its reversed polarity versus styrene monomers.²⁷ That means the unsaturated double bonds of *N*-phenyl maleimide will be in a state of poor-electron and of styrene in a relatively rich-electron state. Then, there was a high tendency for them copolymerized to generate alternating-like St-MI copolymers with the reactivity ratios approaching to zero. The reactivity ratios of styrene/TMMS comonomers, as shown in Table II, were determined by the method of Fineman-Ross,²⁸ Kelen-Tüdös,²⁹ and Tidwell-Mortimer³⁰ while all the instantaneous copolymer yield were under 10%. All the reactivity ratio for the above copolymerization is

Table II. Reactivity ratios of styrene monomers (M_1) and TMMS (M_2)

Monomer (M_1)	Fineman-Ross		Kelen-Tüdös		Tidwell-Mortimer	
	r_1	r_2	r_1	r_2	r_1	r_2
St	0.61 ± 0.04	0.52 ± 0.15	0.59	0.48	0.60	0.49
A-St	0.83 ± 0.02	0.68 ± 0.10	0.82	0.67	0.81	0.66
B-St	0.89 ± 0.02	0.78 ± 0.10	0.88	0.77	0.86	0.76
C-St	0.66 ± 0.03	0.57 ± 0.15	0.65	0.55	0.64	0.54

no longer nearby zero but changed to 0.48–0.88, and is larger than that of the general St-MI copolymerization.^{10–13} This was mainly due to the electron-withdrawing effect of silane side chains to inhibit the carbonyl pairs withdrawing electron effect of double bond within maleimide. Besides, the polarity of styrenes was also affected by their deactivators. The result of the reducing reversed polarity of double bond of TMMS relative to that of styrenes indicated that the above comonomers reaction had changed from an alternating to middle type between alternating and ideal copolymerization. As shown in Table I, all the 4 copolymerization systems, furthermore, exhibit azeotropic points (for example the system of B-St at about $F_1 = f_1 = 0.75$) that represent a composition where polymerization proceeds to relatively high conversion with no change in either feed ratio or copolymer composition.

As shown in Table II, the reactivity orders of TMMS-terminated radical towards the 4 series of styrene monomers was St > C-St > A-St > B-St, indicating that these copolymerizations were also affected by the properties of different side chains within styrene monomers. Since the deactivity effect of ester groups were higher than chlorine or hydrogen atoms, all the molecular weight, PDI or comonomer's reactivity ratio should, therefore, be affected by both steric hindrance and polarity factors of TMMS side chains and the deactivators of styrene simultaneously.

Thermal Properties of Copolymers

Glass Transition Temperature. Glass transition temperature (T_g) of copolymers is one of the important physical properties that may be affected by their composition and morphology. All the T_g data of the 4 series of copolymers are given in Table III. Since the molecular weights of all the 4 series of copolymers were higher than 5.2×10^4 (see Table I), the effect of molecular weight on T_g was not significant. The DSC measurement for all the copolymers showed that all St-MI copolymers had a single T_g that would increase along with the increasing maleimide content. This means that maleimide-segments and styrene-segments within the copolymers should be completely compatible. Theoretically, T_g of the copolymers should comply with the Fox equation:³¹

Table III. Glass transition temperatures of copolymers prepared from styrene monomers (M_1) and TMMS (M_2)

Copolymer	f_2^a	W_2^b	T_g (°C) ^c	T_g (K)	T_{g12} (K) ^d		
MI/St	0.00	0.00	98	371	366		
	0.14	0.29	126	399			
	0.27	0.48	139	412			
	0.43	0.65	142	415			
	0.46	0.68	144	417			
	0.56	0.76	147	420			
	0.69	0.85	153	426			
	0.82	0.92	161	434			
	1.00	1.00	172	445 ^e			
	MI/A-St	0.00	0.00	144		417	390
		0.12	0.18	151		424	
		0.26	0.36	155		428	
		0.36	0.48	157		430	
0.48		0.60	159	432			
0.58		0.69	161	434			
MI/B-St	0.72	0.81	164	437	404		
	0.86	0.91	167	440			
	0.00	0.00	156	429			
	0.11	0.14	160	433			
	0.25	0.30	161	434			
	0.38	0.44	162	435			
	0.49	0.55	163	436			
	0.59	0.65	165	438			
	0.73	0.78	167	440			
	0.87	0.89	169	442			
MI/C-St	0.00	0.00	106	379	368		
	0.13	0.22	124	397			
	0.27	0.41	137	410			
	0.44	0.59	141	414			
	0.47	0.63	143	416			
	0.57	0.71	147	420			
	0.71	0.82	153	426			
	0.83	0.92	161	434			

^aMole fraction of M_2 within the copolymer. ^bWeight fraction of M_2 within the copolymer. ^cThe glass transition temperature (T_g) is analyzed by DSC at a heating rate of 10 °C/min in N_2 . ^dGlass transition temperature calculated by the modified equation of Johnston. ^eThe data of homopolymer were used from the previous study (ref 26).

$$\frac{1}{T_g} = \frac{W_1}{T_{g1}} + \frac{W_2}{T_{g2}} \quad (1)$$

where W_i was the weight fraction of comonomer i within the copolymer and T_{gi} was the homopolymer's

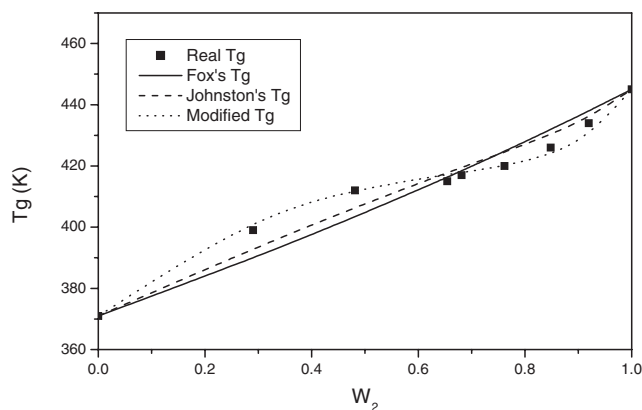


Figure 2. Variation of T_g as a function of the weight fraction of TMMS (W_2) for the St-MI copolymer system.

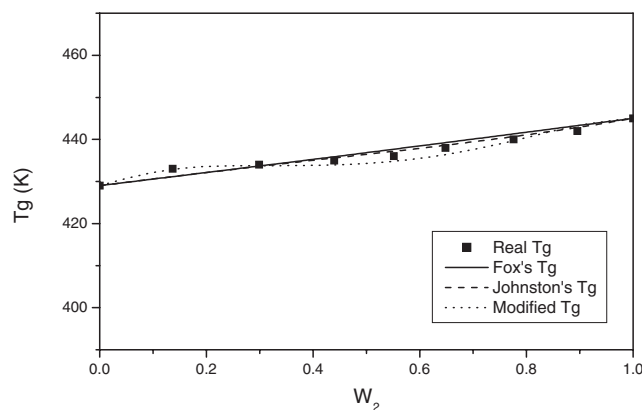


Figure 4. Variation of T_g as a function of the weight fraction of TMMS (W_2) for the B-St/MI copolymer system.

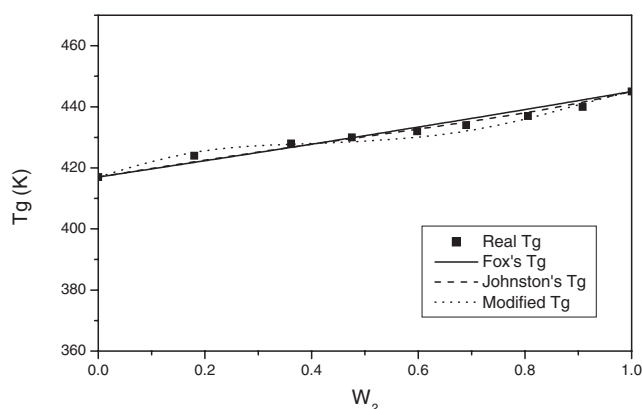


Figure 3. Variation of T_g as a function of the weight fraction of TMMS (W_2) for the A-St/MI copolymer system.

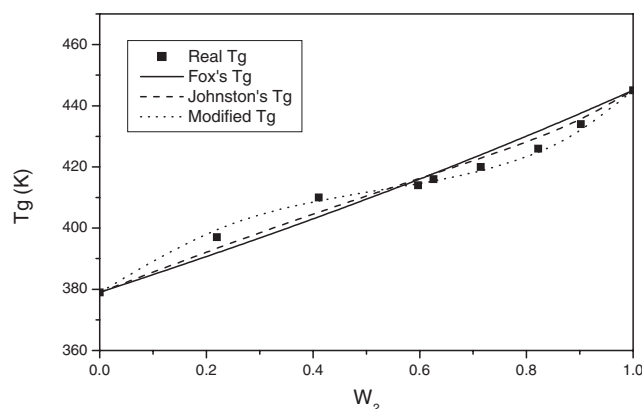


Figure 5. Variation of T_g as a function of the weight fraction of TMMS (W_2) for the C-St/MI copolymer system.

T_g of comonomer i .

By substituting T_{gi} into eq 1, the real T_g s of copolymers matching the Fox theory could be obtained. The relationships between the Fox's T_g -curve and actual T_g 's values, measured by DSC, versus the copolymer compositions are shown in Figures 2–5. Considering the segment structure of TMMS, although its silicon-containing side chain was high mobile segment, molecular motions of these segments could only proceed under a relatively higher temperature or energy due to the high rigidity of the imide structure. As shown in Table III, the T_g s of TMMS homopolymer and all the copolymers were higher than that of polystyrene series, indicating that the rigid degree of the former was higher than that of the latter. In Figures 2–5, the actual T_g curves of copolymers in comparison with the Fox's curves exhibited the positive/negative deviations. The S-like curve of T_g in blending polymer system had also been studied in literature.³² The T_g of the alternating segments within copolymer (T_{g12}) should affect the T_g of the overall copolymer significantly. Besides, each of the 4 series copolymers had a T_g value conformed to match the Fox theory (the

value at the intersection with Fox's curve), postulating that the comonomer segment-distribution for copolymer in azeotropic composition was more random than that in the others leading to offset the alternate segment effect. On the other hand, except the copolymer with azeotropic composition, the molecular motion of the others was easily affected by its alternate segment. Therefore, a higher content of silicon-containing maleimide segment within copolymer would produce a negative deviation of T_g due to the effect of styrene segment, and on the contrary, a lower content would produce a positive deviation of T_g due to the effect of maleimide segment. In order to understand the effect of the alternate segment for the actual T_g s of copolymers, the Johnston's equation³³ modified from Fox theory was applied:

$$\frac{1}{T_g} = \frac{W_1 P_{11}}{T_{g11}} + \frac{W_2 P_{22}}{T_{g22}} + \frac{W_1 P_{12} + W_2 P_{21}}{T_{g12}} \quad (2)$$

where P_{ij} : probability of forming ij -segment, T_{gii} : homopolymer's T_g of comonomer i , and T_{gij} : glass transition temperature of alternating segment within copolymer. The P_{ij} could be calculated as:

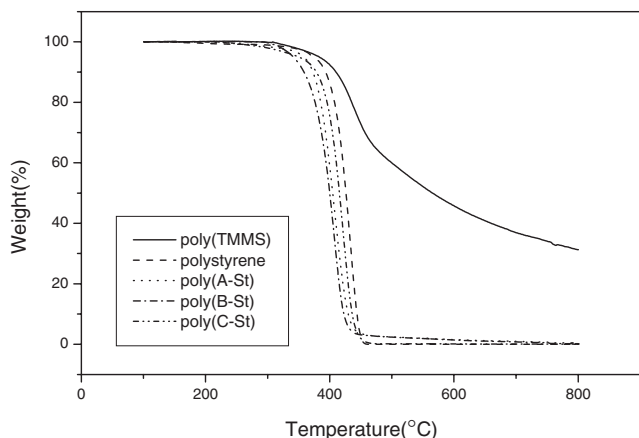


Figure 6. TGA traces of poly(TMMS) and polystyrene series homopolymers in N_2 .

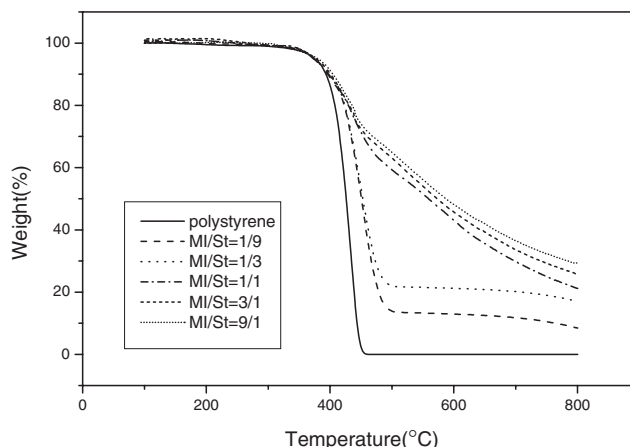


Figure 8. TGA traces of Styrene-TMMS copolymers in N_2 .

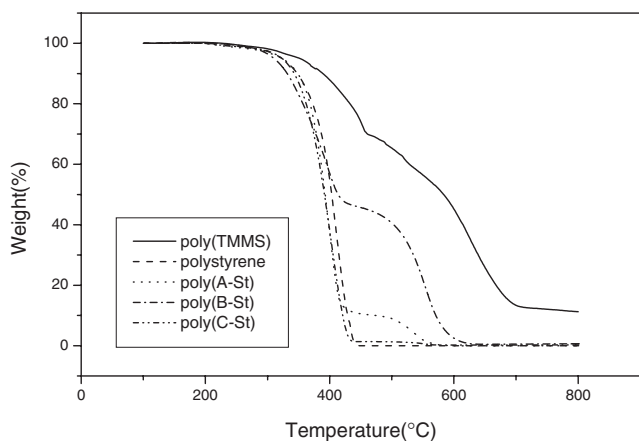


Figure 7. TGA traces of poly(TMMS) and polystyrene series homopolymers in air.

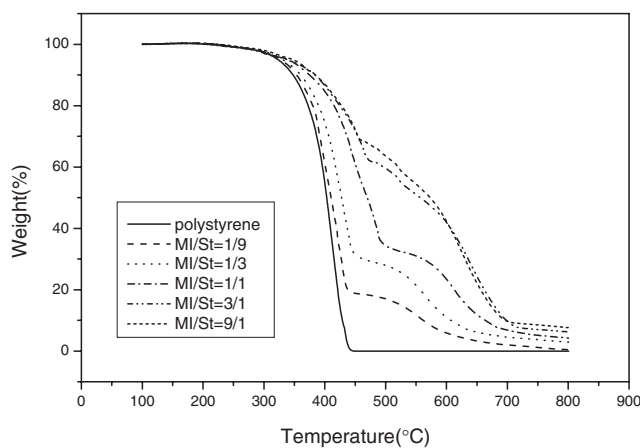


Figure 9. TGA traces of Styrene-TMMS copolymers in air.

$$P_{12} = \frac{1}{1 + r_1 \frac{[1]}{[2]}} = \frac{1}{1 + r_1 \frac{f_1}{f_2}}$$

$$P_{11} = 1 - P_{12}$$

$$P_{21} = \frac{1}{1 + r_2 \frac{[2]}{[1]}} = \frac{1}{1 + r_2 \frac{f_2}{f_1}}$$

$$P_{22} = 1 - P_{21}$$

It could also be noticed in Figures 2–5 that the Johnston's curves of copolymers did not match the actual T_g curves. That means the effect of its T_{g12} on the real T_g of copolymer cannot be sufficiently described by Johnston's equation. Considering the influence of T_{g12} that should be more significant, a modified Johnston's equation can be taken as following:³⁴

$$\frac{1}{T_g} = \frac{W_1 P_{11}}{T_{g11}} + \frac{W_2 P_{22}}{T_{g22}} + \frac{W_1^\alpha P_{12} + W_2^\alpha P_{21}}{T_{g12}} \quad (3)$$

where α is a parameter according to the effect of alter-

nating segment. Substitution of actual T_g of all the 4 series copolymers into the eq 3 gave an average of T_{g12} so as to obtain the modified T_g - W_2 curves. The value α was determined by a try-and-error method so as to make the modified T_g - W_2 curves approaching to the actual T_g as possible as it can be. It could also be noticed from Figures 2–5 that the actual T_g of all the 4 series of copolymers were proved to match our modified curves while $\alpha = 1.24$ in St, $\alpha = 1.14$ in A-St, $\alpha = 1.10$ in B-St, and $\alpha = 1.22$ in C-St series. That means the T_g of these copolymers could be sufficiently described by our modified equation as a result of increasing the affect of weight ratio on T_{g12} . Besides, the S-shaped curve of deviation is more not obvious as the difference of T_g between comonomers is small like B-St system.

Thermal Stability

Thermal properties of 4 series St-MI copolymers and homopolymers measured by TGA in N_2 or air as shown in Figures 6–9 are also listed in Table IV. As seen in Figure 6 and Table IV, $T_{5\%}$ of all the sty-

Table IV. Thermal properties of copolymers prepared from styrene monomers (M_1) and TMMS (M_2)

Monomer (M_1)	F_1^a	f_1^b	Si (%) ^c	5 wt % Loss (°C) ^d		Char at 800 °C (%) ^d	
				N ₂	Air	N ₂	Air
St	1.00	1.00	0.0	373	323	0.0	0.0
	0.90	0.86	3.1	374	327	8.4	0.5
	0.75	0.73	5.2	375	332	17.2	2.9
	0.50	0.54	7.3	376	338	21.1	4.2
	0.25	0.31	9.0	378	342	25.8	6.2
	0.10	0.18	9.9	380	348	29.3	7.7
	0.00	0.00	10.7	384	352	31.3	12.0
A-St	1.00	1.00	0.0	345	323	0.2	0.0
	0.90	0.88	1.9	348	326	8.5	0.8
	0.75	0.74	3.9	355	330	18.1	2.6
	0.50	0.52	6.4	363	338	22.4	4.5
	0.25	0.28	8.6	374	344	26.5	6.8
	0.10	0.14	9.7	381	349	29.2	8.1
	0.00	0.00	10.7	384	352	31.3	12.0
B-St	1.00	1.00	0.0	341	320	0.4	0.2
	0.90	0.89	1.5	344	322	6.2	0.4
	0.75	0.75	3.2	351	327	15.8	2.3
	0.50	0.51	5.9	361	335	21.7	4.0
	0.25	0.27	8.3	372	343	23.6	6.2
	0.10	0.13	9.6	379	348	26.1	7.8
	0.00	0.00	10.7	384	352	31.3	12.0
C-St	1.00	1.00	0.0	350	323	0.0	0.3
	0.90	0.87	2.4	354	326	8.1	0.5
	0.75	0.73	4.4	359	331	17.0	2.2
	0.50	0.53	6.7	368	339	20.9	4.1
	0.25	0.29	8.8	378	346	25.5	6.0
	0.10	0.17	9.7	381	350	29.0	7.1
	0.00	0.00	10.7	384	352	31.3	12.0

^aMole fraction of M_1 in the feed. ^bMole fraction of M_1 within the copolymer. ^cThe theoretical calculation values of Si composition ratio. ^dThermal properties are analyzed by TGA at a heating rate of 20 °C/min in N₂ or in air. ^eHomopolymer of St, A-St, B-St and C-St with a M_n about 1.5×10^5 , 5×10^4 , 4.5×10^4 and 7.5×10^4 .

rene series homopolymers was higher than 341 °C when were heated under nitrogen gas. The A-St and B-St homopolymers exhibited a lower $T_{5\%}$ due to the decomposition reactions of their ester side-chains. Decomposition reactions of TMMS homopolymer should be initialized mainly by scission of silane side-chains because of owning lower bond energy,²⁶ and further produced scission of the maleimide main-chain along with carbonization to form char yield of solid residue. Thermal stability of all the homopolymers in N₂ was higher than that in air because of the oxidization reaction. When all homopolymers were heated in air, as shown in Figure 7, decomposition of TMMS, A-St and B-St series changed from one-stage to two-stage. The decomposition of the second stage should be the oxidized combustion reaction of ester, silane or imide fragment under the high temperature.

The TGA traces of St-MI copolymers measured in nitrogen and air are shown in Figures 8 and 9, respectively. The scission of all the copolymers in nitrogen

and air was a 1-stage and a 2-stage process, respectively. Decomposition reactions in N₂, as shown in Figure 8, should be initialized mainly by scission of silane side-chains, and further produced scission of the maleimide and styrene main-chain along with carbonization to form char yield of solid residue. Thermal stability of all the St-MI copolymers in N₂ was higher than that in air because of the oxidization reaction. When all copolymers were heated in air, decomposition, as shown in Figure 9, changed from one-stage to two-stage. The decomposition of the second stage should be the oxidized combustion reaction under the high temperature. The char yield of the silicon-containing copolymer pyrolyzed in air was generally increased with increasing the silicon content. This might due to the high oxidation resistance of silicon-containing solid residue, such as silicon oxide or silicon-carbon alloy.³⁵ Besides, maximum rates of weight loss of these polymers were decreased and the pyrolysis curves moved to higher temperature as the silicon content was increased. The results of high-

er char yield and lower weight loss rates during pyrolysis of silicon-containing system showed that the introduction of silane side-chain into polymers improved flame retardancy.

The TGA traces of 4 styrene series copolymers measured in nitrogen and air are alike. As comparing to the $T_{5\%}$ in Table IV, the thermal stability of each copolymer series was higher than that of styrene-series homopolymers. The initial temperature of pyrolysis of all the 4 series copolymer was generally increased with increasing the TMMS content. This means that the silicon-containing side chains could increase thermal stability of copolymers in nitrogen or in air. Also, the char yield of all the 4 series copolymer was generally increased with increasing the silicon content both in N_2 and in air environment. The higher char yield and lower weight loss rates showed that the introduction of silane side-chain into polymers improved flame retardancy.

CONCLUSION

The radical copolymerizations of 4 styrene-series with silicon-containing maleimide were carried out to get some novel copolymers. The factors, such as steric hinderance and chain transfer caused by silicon-containing maleimide, could affect the molecular weight, polydispersity index and segments distribution of copolymer. The actual T_g 's' curves versus the composition ratios for all the above copolymers were very consistent with the modified Johnston's equation as increasing the weight ratio of alternate segments. The maleimide-segments and styrene-segments within these copolymers were completely compatible and the thermal stability and flame retardancy of polystyrene could be enhanced simultaneously *via* the introduction of silicon-containing maleimide.

Acknowledgment. The author would like to acknowledge the financial support for this research of the National Science Council, Taiwan, under the contact NSC 94-2622-E-233-001.

REFERENCES

1. D. C. Wu, C. Y. Hong, C. Y. Pan, and W. D. He, *Polym. Int.*, **52**, 98 (2003).
2. F. Yilmaz, L. Cianga, Y. Guner, L. Topppare, and Y. Yagci, *Polymer*, **45**, 5765 (2004).
3. C. Soykan and I. Erol, *J. Appl. Polym. Sci.*, **91**, 964 (2004).
4. G. Liu, L. Zhang, C. Gao, and X. Qu, *J. Appl. Polym. Sci.*, **98**, 1932 (2005).
5. S. L. Oswal, N. S. Sarkar, V. K. Bhandari, H. B. Oza, and C. B. Patel, *Iranian Polym. J.*, **13**, 297 (2004).
6. C. B. Patel, N. I. Malek, and S. L. Oswal, *J. Macromol. Sci., Part A: Pure Appl. Chem.*, **43**, 289 (2006).
7. T. Cakir, I. E. Serhatli, and A. Onen, *J. Appl. Polym. Sci.*, **99**, 1993 (2006).
8. W. Y. Chiang and F. C. Ding, *J. Polym. Res.*, **7**, 251 (2000).
9. I. A. Salman, Al-Sagheer, A. Fakhria, and M. Z. Elsabee, *J. Macromol. Sci., Part A: Pure Appl. Chem.*, **34**, 1207 (1997).
10. J. A. Seiner and M. Litt, *Macromolecules*, **4**, 308 (1971).
11. R. E. Cais, R. G. Farmer, D. J. T. Hill, and J. H. O'Donnell, *Macromolecules*, **12**, 835 (1979).
12. M. Wang, X. Zhu, and L. Zhang, *J. Appl. Polym. Sci.*, **75**, 267 (2000).
13. M. Switala-Zeliazkow, *Polym. Degrad. Stab.*, **74**, 579 (2001).
14. W. Y. Chiang and J. Y. Lu, *J. Appl. Polym. Sci.*, **50**, 1007 (1993).
15. K. D. Ahn, J. S. Koo, and C. M. Chung, *J. Polym. Sci., Part A: Polym. Chem.*, **34**, 183 (1996).
16. T. Iijima, M. Hirano, W. Fukuda, and M. Tomoi, *Eur. Polym. J.*, **29**, 1399 (1993).
17. M. N. Teerenstra, D. R. Suwier, B. Van Mele, L. Teuwen, M. Maassen, H. J. Van Den Berg, and C. E. Koning, *J. Polym. Sci., Part A: Polym. Chem.*, **38**, 3550 (2000).
18. D. R. Suwier, P. A. M. Steeman, M. N. Teerenstra, M. A. J. Schellekens, B. Vanhaecht, M. J. Monteiro, and C. E. Koning, *Macromolecules*, **35**, 6210 (2002).
19. S. T. Kim, J. B. Kim, C. M. Chung, and K. D. Ahn, *J. Appl. Polym. Sci.*, **66**, 2507 (1997).
20. K. D. Ahn, D. I. Koo, and C. G. Willson, *Polymer*, **36**, 2621 (1995).
21. K. D. Ahn, C. M. Chung, H. S. Jo, and J. M. Rhee, *Polym. Int.*, **47**, 407 (1998).
22. W. Y. Chiang and Y. C. Lin, *J. Appl. Polym. Sci.*, **83**, 2791 (2002).
23. W. Y. Chiang and M. L. Lee, *J. Appl. Polym. Sci.*, **90**, 1032 (2003).
24. G. N. Taylor, T. M. Wolf, and L. E. Stillwagon, *Solid State Technol.*, **27**, 145 (1984).
25. C. X. Chen, R. J. Hurditch, D. W. Johnson, and D. J. Nawrocki, *Proceedings of SPIE, The International Society for Optical Engineering*, **4690 I**, 262 (2002).
26. W. J. Shu, J. C. Ho, and L. H. Perng, *Eur. Polym. J.*, **41**, 149 (2005).
27. S. R. Turner, R. A. Arcus, C. G. Houle, and W. R. Schleigh, *Polym. Eng. Sci.*, **26**, 1096 (1986).
28. M. Fineman and S. D. Ross, *J. Polym. Sci.*, **2**, 259 (1950).
29. T. Kelen and F. Tudos, *J. Macromol. Sci. Pt. A, Chem.*, **9**, 1 (1975).
30. P. W. Tidwell and G. A. Mortimer, *J. Polym. Sci., Part A*, **3**, 369 (1965).
31. T. G. Fox, *Bull. Am. Phys. Soc.*, **1**, 123 (1956).
32. T. K. Kwei, *J. Polym. Sci., Polym. Lett. Ed.*, **22**, 307 (1984).
33. N. W. Johnston, *J. Macromol. Sci., Rev. Macromol. Chem.*, **C14**, 215 (1976).
34. W. J. Shu, L. H. Perng, W. K. Chin, and C. Y. Hsu, *J. Macromol. Sci., Part A: Pure Appl. Chem.*, **40**, 897 (2003).
35. M. Foster, B. Darlington, J. Scharff, and A. Champion, *Surf. Sci.*, **375**, 35 (1997).